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GRADE 12
PHYSICAL SCIENCES: PAPER 2
JULY PRELIMINARY EXAM
2024



Examiner : R. Stroud
Moderator : Y. Patterson

Time: 3 hours
Marks : 200

Name : _____ **Memorandum** _____ = _____ %
200

PLEASE READ THE FOLLOWING INSTRUCTIONS CAREFULLY

1. This question paper consists of 26 pages and a Data Sheet. Please check that your paper is complete.
2. Answer ALL the questions in the spaces provided on the paper.
3. Read the questions carefully and answer what is asked.
4. Use the data and formulae whenever necessary.
5. Show your workings in all calculations.
6. Units need not be included in the workings of your calculations, but appropriate units should be shown in the answer.
7. Answers must be expressed in decimal format, not left as proper fractions.
8. Where appropriate, express answers to TWO decimal places.
9. It is in your own interest to write legibly and present your work neatly.
10. Blank spaces are provided between questions and at the end of the paper if more space is required. Clearly indicate the question number of your answer should you use this extra space.

Name : _____

Question	Possible Mark	Mark Achieved
1	20	
2	17	
3	27	
4	30	
5	29	
6	17	
7	28	
8	16	
9	16	
Total	200	

QUESTION 1 MULTIPLE CHOICE

Answer these questions on the multiple choice answer sheet given below. Make a cross (X) in the box corresponding to the letter you consider to be correct.

eg	A	B	~~X~~	D
1.1	A	B	C	D
1.2	A	B	C	D
1.3	A	B	C	D
1.4	A	B	C	D
1.5	A	B	C	D
1.6	A	B	C	D
1.7	A	B	C	D
1.8	A	B	C	D
1.9	A	B	C	D
1.10	A	B	C	D

Here C is chosen as an example

[20]

Question 1 Multiple Choice

1.1 One mole of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ contains:

- A $6,02 \times 10^{23}$ atoms
- B 5 water molecules
- C 4 mol oxygen atoms
- D 5 mol water molecules

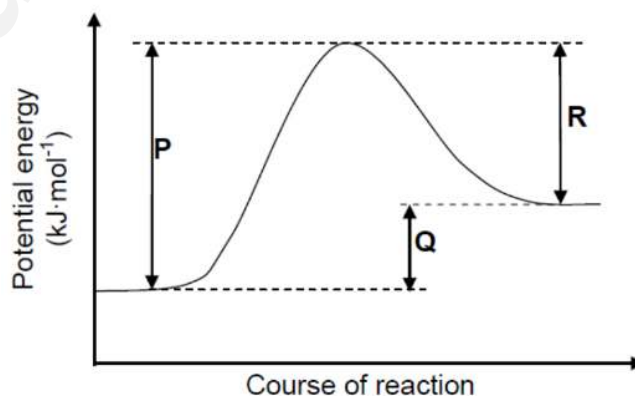
1.2 The concentration of a $0,2 \text{ dm}^3$ solution $\text{Mg}(\text{OH})_2$ is $0,8 \text{ mol} \cdot \text{dm}^{-3}$. Assuming complete dissociation, the mole number of hydroxide ions is:

- A 0,16
- B 4,0
- C 0,32
- D 8,0

1.3 Which row correctly shows the dominant intermolecular forces in the compounds listed?

	BH_3	H_2S	HCN
A	Dipole-dipole	Hydrogen bonding	Hydrogen bonding
B	London forces	Dipole-dipole	Dipole-dipole
C	Hydrogen bonding	London forces	Hydrogen bonding
D	London forces	Hydrogen bonding	Dipole-dipole

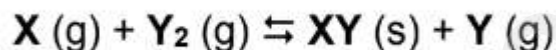
1.4 The energy profile of a reversible reaction is given below.



The best description of the changes in **P**, **Q** and **R** with the **removal** of a catalyst for the forward reaction is:

	P	Q	R
A	increases	stays the same	increases
B	decreases	stays the same	decreases
C	decreases	decreases	decreases
D	increases	increases	increases

- 1.5 1 mol of **X** and 1 mol of **Y₂** are placed in a sealed container with a volume of 1 dm³. **X** and **Y₂** react according to the equation given below:



The equilibrium constant is determined to be $K_c = 2$ at a certain temperature. The concentrations of the substances at equilibrium are:

- A $[\text{X}][\text{Y}_2] = 0,5[\text{Y}]$
- B $2[\text{Y}] = [\text{X}][\text{Y}_2]$
- C $[\text{X}][\text{Y}] = [\text{XY}]$
- D $[\text{X}] = 2[\text{Y}]$

- 1.6 Which of the following acidic solutions, each with a concentration of 1 mol.dm⁻³, will have the lowest conductivity?

- A H₂SO₄ (aq)
- B HNO₃ (aq)
- C H₂SO₃ (aq)
- D HCl (aq)

- 1.7 Which of the following solutions of ions will oxidise Zn, but will not oxidise Sn under standard conditions?

- A Cu²⁺ (aq)
- B Mg²⁺ (aq)
- C Fe³⁺(aq)
- D Fe²⁺ (aq)

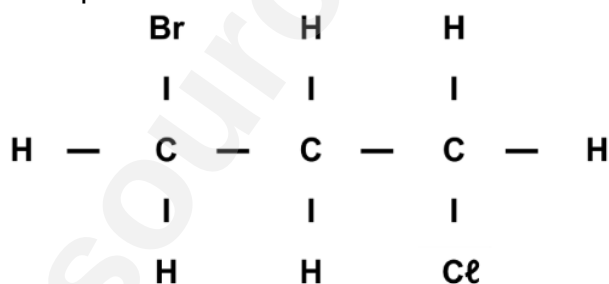
1.8 Which reaction correctly describes the reduction half reaction during the electrolysis of brine using the membrane cell?

- A $\text{Na}^+ + \text{e} \rightarrow \text{Na}$
- B $2\text{Cl}^- + 2\text{e} \rightarrow \text{Cl}_2$
- C $2\text{H}_2\text{O} + 2\text{e} \rightarrow \text{H}_2 + 2\text{OH}^-$
- D $2\text{H}^+ + 2\text{e} \rightarrow \text{H}_2$

1.9 Which of the following is an unsaturated hydrocarbon?

- A C_2H_6
- B C_4H_6
- C CH_3COOH
- D CH_4

1.10 Given the organic compound below:



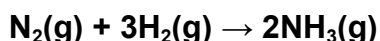
The correct IUPAC name for a **positional isomer** of the compound above is:

- A 1-bromo-2-chloropropane
- B 3-bromo-1-chloropropane
- C 2-bromo-3-chloropropane
- D 1-bromo-3-chloropropane

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Question 2

The reaction below is used in the Haber process to manufacture ammonia.



The boiling points of the substances in the reaction are as follows:

Substance	Boiling Point (°C)
H ₂	-253
N ₂	-196
NH ₃	-33

- 2.1 Explain an *intramolecular bond*. (1)

A bond that occurs between atoms within molecules. ✓

- 2.2 Name the specific intramolecular bond that occurs in NH₃. (2)

Polar ✓ covalent bond ✓

- 2.3 Refer to the intermolecular forces and explain the difference in boiling point between NH₃ and N₂. (3)

N₂ is a nonpolar molecule forming weak London forces ✓ between molecules.

NH₃ forms stronger hydrogen bond forces ✓ between molecules and London forces. (HBF are the strongest van der Waals forces)

More energy is needed to overcome the hydrogen bond forces than London forces, so the BP for NH₃ is higher. ✓

- 2.4 Refer to the type and relative strength of the intermolecular forces, to explain why H₂ has a lower boiling point than N₂. (2)

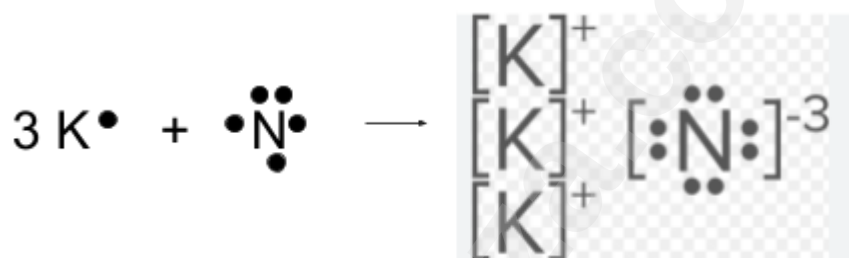
Both have London forces, ✓ but as H₂ is a smaller molecule, ✓ with less electrons, less London forces form and are easier to overcome than for N₂.

- 2.5 Explain why the bonding between a metal and a nonmetal can be seen as an intermolecular force. (2)

It is a weak force of attraction ✓ between ions ✓ / transfer of e- to form ions with coulombic forces of attraction.

- 2.6 Potassium nitride has a boiling point of 537 °C.

- 2.6.1 Using Lewis diagrams, show the bonding of the atoms to form the salt K₃N. (3)



- 2.6.2 Compare the boiling points of NH₃ and K₃N. Using your knowledge of intermolecular forces, explain this difference. (4)

NH₃ forms hydrogen bond forces between molecules. ✓

K₃N has ionic electrostatic forces between ions. ✓

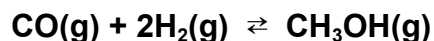
The electrostatic forces are much stronger than hydrogen bond forces, ✓

Needing more energy to overcome the forces, increasing the BP. ✓

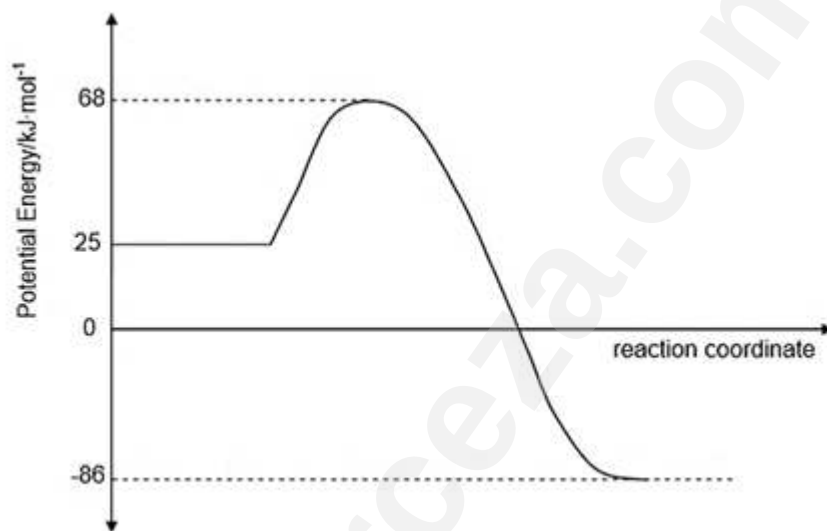
[17]

Question 3

The reaction below between carbon monoxide and hydrogen can be used industrially to make methanol. The process is carried out at 400 °C and pressure between 5000 kPa and 10 000 kPa.



The energy profile below shows the energy change during this reaction:



- 3.1 Calculate the enthalpy change (ΔH) of this reaction. (2)

$$\begin{aligned} \Delta H &= E_P - E_R & \text{OR} & & \Delta H &= E_{in} - E_{out} \text{ (usually for bond E)} \\ &= -86 - 25 \checkmark & & & &= (43) - (154) \\ &= -111 \text{ kJ.mol}^{-1} \checkmark & & & & \end{aligned}$$

- 3.2 State whether this is an endothermic or exothermic reaction. (1)

Exothermic ✓

- 3.3 State the energy at which the activated complex is formed. (1)

68 KJ.mol⁻¹ ✓

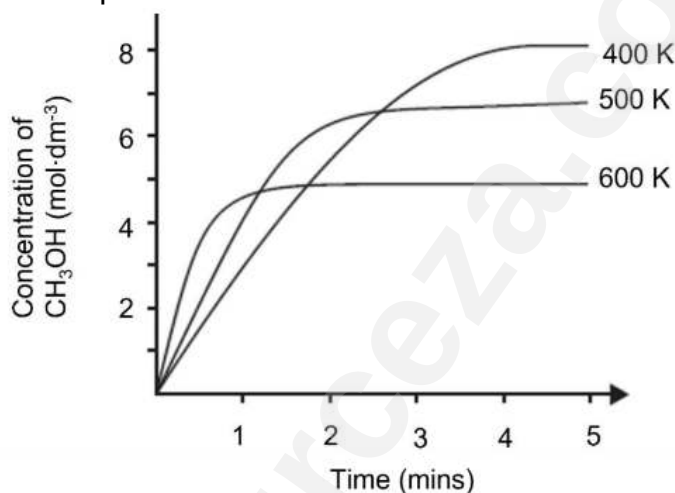
- 3.4 Define the term *activated complex*. (2)

A high energy, unstable, transition state between reactants and products. ✓✓

- 3.5 Calculate the activation energy for the **reverse** reaction. (2)

Reverse reaction: $E_{in} = 68 - (-86) ✓$
 $= 154 \text{ kJ.mol}^{-1} ✓$

- 3.6 CO and H₂ are placed in a closed container. The graph below shows how the concentration of methanol produced changes with different temperatures. The pressure remained constant.



- 3.6.1 Calculate the average rate of production of methanol in the first 4 minutes at 400 K. (3)

$$\text{Rate} = \frac{\Delta[\text{CH}_3\text{OH}]}{t} ✓$$

$$= \frac{8-0}{4-0} ✓$$

$$= 2 \text{ mol.dm}^{-3}.\text{min}^{-1} ✓ \quad 0,033 \text{ mol.dm}^{-3}.\text{s}^{-1}$$

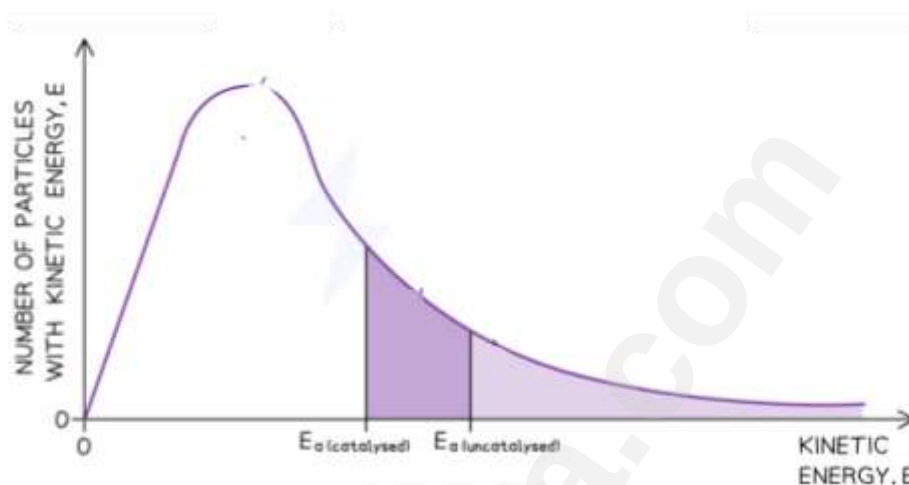
- 3.6.2 State at which temperature the rate was the highest. (1)

600K ✓

- 3.6.3 Describe ONE effect, other than an increased rate, that the increase in temperature would have on the reaction. (2)

The equilibrium concentration of methanol decreases. ✓ ✓ K_c decreases

- 3.7 The Maxwell-Boltzman curve below shows the effect of a reaction with or without a catalyst.



Using the collision theory, explain how the rate of the reaction is affected when a catalyst is added. (4)

E_a is lowered with a catalyst. ✓ Each reaction requires less activation energy. ✓

The catalyst also provides an alternate pathway assisting with the correct orientation of the reactants, while remaining unchanged.

With more energy ✓ and more effective collisions made per unit time, ✓ the rate increases.

- 3.7 61,6 g carbon monoxide and 10 g hydrogen are placed in a closed container at STP and allowed to react completely. Determine the limiting reagent. (6)

$$n(\text{CO}) = \frac{m}{M}$$

$$= \frac{61,6}{28} \quad \checkmark$$

$$= 2,2 \text{ mol}$$

$$n(\text{H}_2) = \frac{m}{M}$$

$$= \frac{10}{2} \quad \checkmark$$

$$= 5 \text{ mol}$$

$$1 : 2 \quad \checkmark \quad \text{OR} \quad 1 : 2 \quad 1 : 2$$

$$\frac{2,2}{1} = 2,2 : \frac{5}{2} = 2,5 \quad \checkmark \quad \frac{1}{2} (5) : 5 \quad 2,2 : 2(2,2) = 4,4$$

$$2,2 < 2,5 \quad \checkmark \quad = 2,5 \quad \text{enough H}_2$$

CO limits ✓ not enough CO
(State the obvious, follow logical order)

- 3.8 Calculate the mass of methanol gas that would be produced should the reaction run to completion with the quantities given in 3.7. (3)

CO : CH₃OH (Only limiting values can be used across the equation)

1 : 1

2,2 : 2,2

m (CH₃OH) = nM

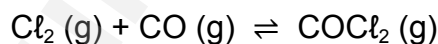
= (2,2)(32)

= 70,4 g

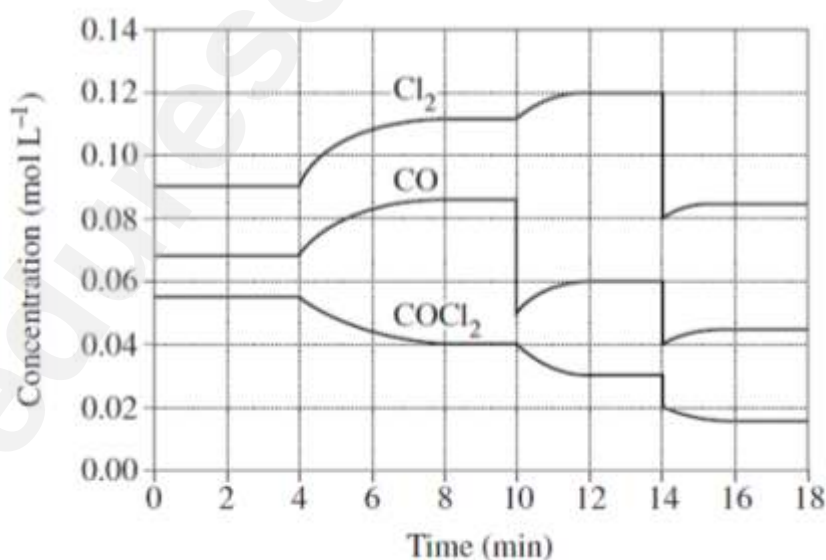
[27]

Question 4

- 4.1 Consider the following equilibrium reaction:



The graph shows a series of disturbances of this equilibrium.



- 4.1.1 State Le Chatelier's Principle. (3)

When an external stress (change in temperature, pressure or concentration) ✓
is applied to a system in dynamic chemical equilibrium, ✓ the equilibrium will

change in such a way as to counteract that stress. ✓

- 4.1.2 The first change at $t = 4$ minutes is caused by an increase in temperature. Use Le Chatelier's Principle to determine whether the forward reaction is endothermic or exothermic. (4)

According to Le Chatelier, an increase in temperature favours the reaction that uses heat - the endothermic reaction. ✓

The increase in temperature produced more reactants, ✓ thus favouring the reverse reaction. ✓

The forward reaction is thus exothermic. ✓

- 4.1.3 At what time is equilibrium established after this change in temperature? (1)

8 min ✓

- 4.1.4 Consider the disturbances at $t = 10$ min and $t = 14$ min. Which one of these was caused by a change in pressure? (1)

14 min ✓ (s also accepted due to typo)

- 4.1.5a Determine whether the pressure was increased or decreased. Give a reason for your answer. (2)

There was a sudden drop in concentration of all substances, ✓ so pressure was decreased. ✓

- 4.1.5b Explain, using Le Chatelier's Principle how this change in pressure affected the yield of COCl_2 . (3)

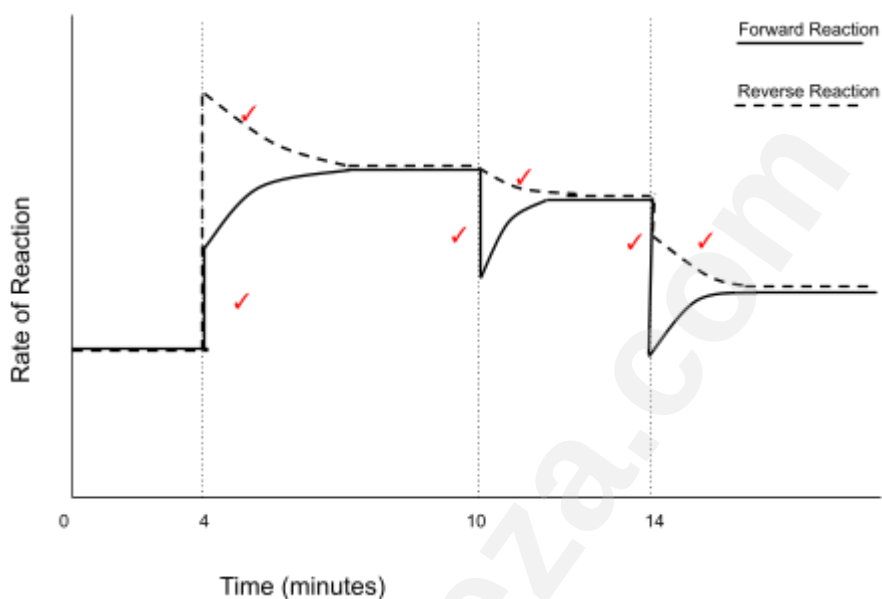
According to Le Chatelier as pressure is decreased, the reaction that will

increase pressure is favoured. ✓

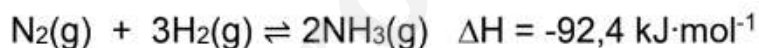
The reverse reaction produces 2 moles to 1, and is thus favoured. ✓

The COCl_2 yield is reduced. ✓

- 4.1.5 Draw a rate of reaction graph to show the changes in rates relative to those given in the concentration graph above. (6)



- 4.2 Working with the Harbor process, where the chemical reaction is:



Solly and Lebo placed 5 g of hydrogen gas in an empty 3 dm³ container with an unknown amount of nitrogen gas (x moles), and sealed the container. The temperature was maintained at 420K and the concentration of ammonia was found to be 0,5 mol.dm⁻³ at equilibrium.

- 4.2.1 Write down the expression for the equilibrium constant for the Harbour process. (2)

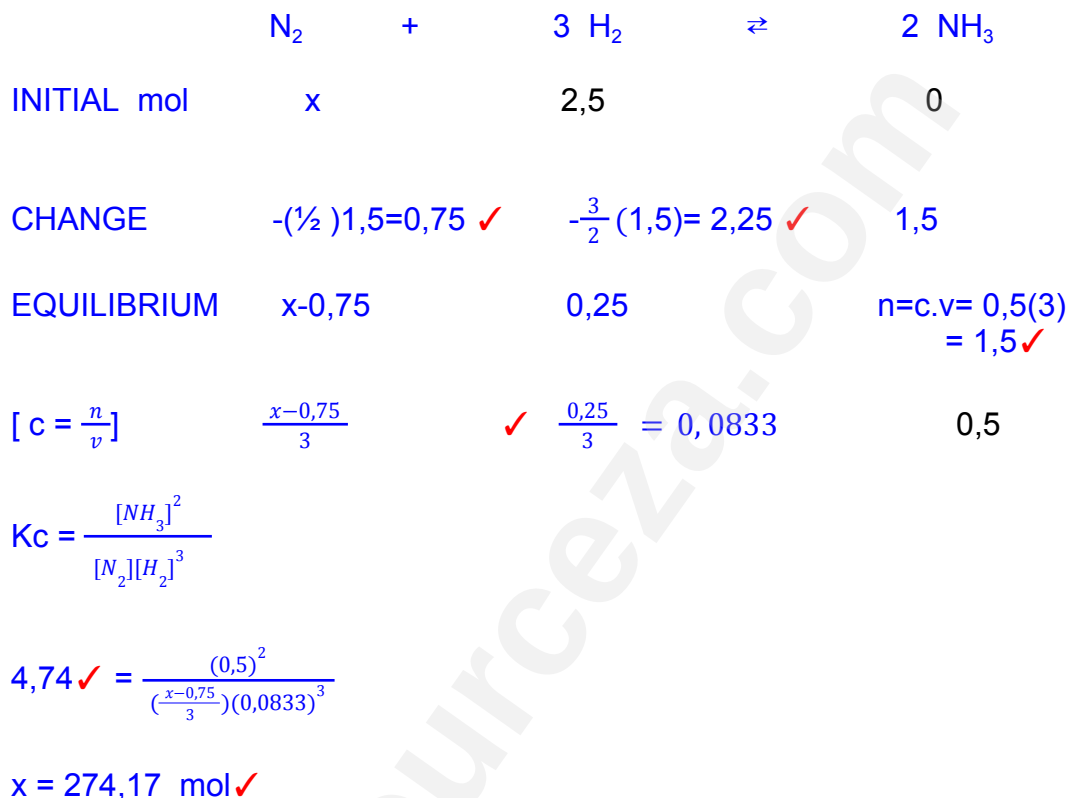
$$K_c = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3} \checkmark \checkmark$$

- 4.2.2 Calculate the number of moles of hydrogen added to the container. (2)

$$n = \frac{m}{M} = \frac{5}{2} \checkmark = 2,5 \text{ mol} \checkmark$$

- 4.2.3 The equilibrium constant for the harbour process at 420 K is given as 4,74. Determine the value of **x moles** of nitrogen originally placed in the container.

(6)



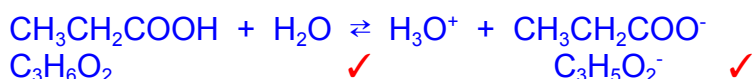
[30]

Question 5

- 5.1 Propanoic acid occurs naturally in Swiss cheese in a concentration that can be as high as 1 %. Sodium and calcium salts are food additives used in processed cheeses to slow down the formation of moulds.

- 5.1.1 Write a balanced reaction equation to show the ionisation of propanoic acid in water.

(2)



- 5.1.2 Propanoic acid is a weak acid. Explain what this means.

(2)

It only partly ionises in an aqueous solution. ✓ ✓

- 5.1.3 Sodium hydroxide reacts with propanoic acid. Write a balanced chemical equation for this reaction. (2)



- 5.1.4 Name the organic product formed in the reaction in question 5.1.3. (1)

Sodium propanoate

- 5.1.5 Define *hydrolysis of a salt*. (1)

The reaction of an ion (from a salt) with water.

- 5.1.6 Using hydrolysis explain whether $\text{CH}_3\text{CH}_2\text{COONa}$ will be an ACIDIC, BASIC or NEUTRAL salt. (4)



Na^+ will remain as ions: from a strong base

$\text{CH}_3\text{CH}_2\text{COO}^-$ from a weak acid will undergo hydrolysis



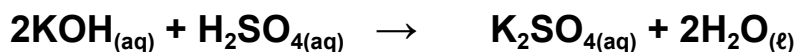
$[\text{H}^+] < [\text{OH}^-]$ so the salt is basic

5.2



During a different titration, a standard solution of sulfuric acid, H_2SO_4 (concentration $0,4 \text{ mol} \cdot \text{dm}^{-3}$), and a potassium hydroxide solution of unknown concentration react.

The chemical equation representing this titration is given below:



5.2.1 Supply labels for the equipment A and B. (2)

A : burette ✓

B: erlenmeyer / conical flask ✓

5.2.2 State where the sulfuric acid should be placed, in A or B. Give a reason for your answer. (2)

In A. ✓ It is the titrant of known concentration. ✓

5.2.3 Assuming the sulfuric acid fully ionises in solution, calculate the concentration of **hydronium** ions in the standard solution of sulfuric acid. (2)

$$c = 0,4(2) \checkmark = 0,8 \text{ mol.dm}^{-3} \checkmark$$

30 cm³ of KOH was neutralised by 42 cm³ of H₂SO₄.

5.2.6 Calculate the concentration of the potassium hydroxide needed to make up the solution used in this titration. (Answer space continues to the next page) (5)

$$\frac{c_{\text{ava}}}{c_{\text{bvb}}} = \frac{n_{\text{a}}}{n_{\text{b}}} \checkmark$$

OR KOH : H₂SO₄

$$\frac{0,4(42)}{c_{\text{b}}(30)} = \frac{1}{2} \checkmark \checkmark \checkmark$$

2 : 1 ✓

$$c_{\text{b}} = \frac{0,4(40)(2)}{30}$$

$$n(\text{H}_2\text{SO}_4) = c.v$$

$$= 1,12 \text{ mol.dm}^{-3} \checkmark$$

$$2(0,0168) = 0,0336 \checkmark$$

$$= 0,4(0,042)$$

$$= 0,0168 \text{ mol} \checkmark$$

$$c(\text{KOH}) = n/v$$

$$= \frac{0,0336}{0,030} \checkmark$$

$$= 1,12 \text{ mol.dm}^{-3} \checkmark$$

5.2.7 Determine the mass of K_2SO_4 produced. (3)

$$m(\text{K}_2\text{SO}_4) = nM \checkmark$$

$$= (0,0168)(174) \checkmark$$

$$= 2,92 \text{ g.} \checkmark$$

5.2.8a Define an *amphiprotic substance*. (2)

A substance that can act either as an acid or a base $\checkmark \checkmark$

(depending on what it is paired with.)

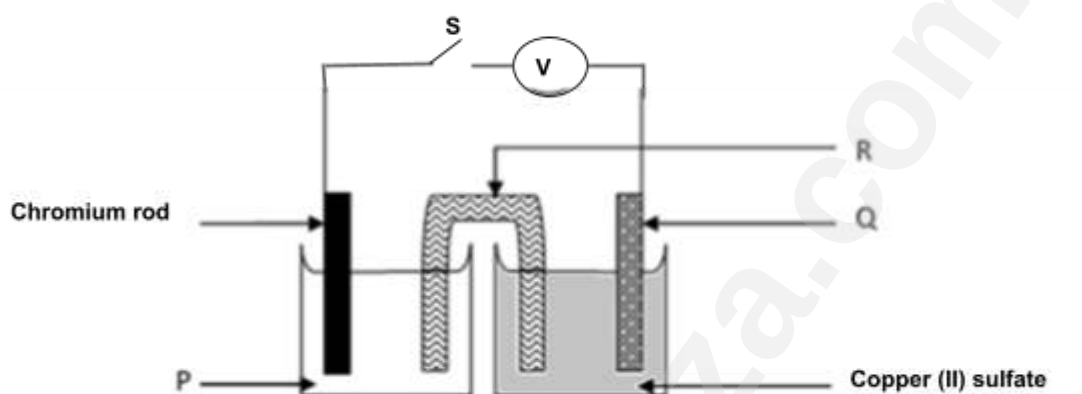
5.2.8b Give an example of an amphiprotic substance from the reaction given in 5.2. (1)



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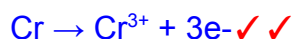
Question 6

An electrochemical cell is constructed consisting of a chromium electrode dipped into a solution containing a suitable electrolyte and an unknown electrode **Q** dipped in a solution of copper(II)sulfate. Study the diagram drawn below and answer the questions which follow.



- 6.1 When the switch **S** is closed, it is found that after some time the chromium electrode corrodes. State whether the chromium electrode is the cathode or anode. Justify your answer with a chemical equation. (3)

Anode ✓



- 6.2 Write the chemical FORMULA for a suitable electrolyte **P**. (1)

$\text{Cr}(\text{NO}_3)_3$ ✓ $\text{Cr}_2(\text{SO}_4)_3$ CrCl_3

- 6.3 Name the **metal** from which electrode **Q** is made. (1)

Copper ✓

- 6.4 Provide the cell notation for this cell. Include standard conditions and phases. (4)

$\text{Cr(s)} / \text{Cr}^{3+}(\text{aq}, 1 \text{ mol.dm}^{-3}) // \text{Cu}^{2+}(\text{aq}, 1 \text{ mol.dm}^{-3}) / \text{Cu(s)} \text{ at } 298 \text{ K}$
 ✓ ✓ ✓ ✓

- 6.5 Calculate the initial emf for the above cell, under standard conditions. (3)

$$\text{Emf} = E_{\text{cat}} - E_{\text{an}} \quad \checkmark$$

$$= +0,34 - (-0,74) \checkmark \text{ (Show the value from the table - 0,74)}$$

$$= 1,08 \text{ V} \checkmark$$

6.6 If the cell has been delivering a current for about 45 minutes, state what will happen in each one of the following cases. State only **INCREASE**, **DECREASE** or **REMAINS UNCHANGED**.

6.6.1 The blue colour of the CuSO_4 solution. (1)

Decrease $\checkmark$

6.6.2 The emf of the cell. (1)

decrease $\checkmark$

6.7 **Describe** and **explain** how electrodes with a larger surface area will influence the voltage (emf) and current delivered by the cell. (2)

Larger surface area will increase the rate of the reaction and allow for more charge to be delivered in a shorter time, so the current will increase $\checkmark$, but not the emf. $\checkmark$ (emf due to difference on table)

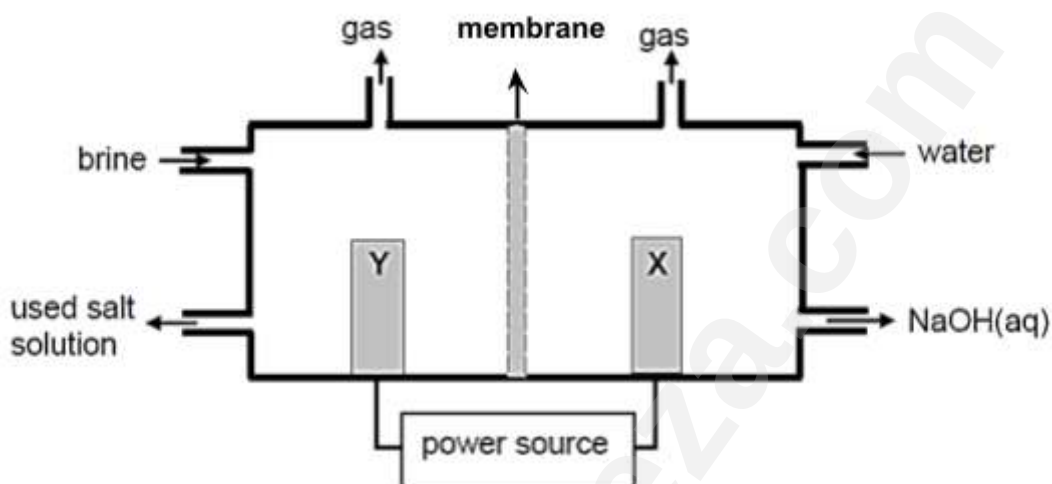
6.8 Name a suitable salt to be used in the salt bridge R. (1)

Potassium nitrate (any gr 1 metal) $\checkmark$

[17]

Question 7

- 7.1 The simplified diagram of a cell used in the chlor-alkali industry is shown below.



- 7.1.1 Describe an *electrolytic cell*. (2)

An electrochemical cell which has electrode reactions sustained by a supply of electrical energy, ✓ converting the electric energy to chemical potential energy. ✓

- 7.1.2 At which electrode, X or Y, is chlorine gas formed? (1)

Y ✓

- 7.1.3 Give the oxidation reaction that occurs in this cell. (2)



- 7.1.4 Explain the function of the membrane. (2)

It is a cationic membrane only allowing the Na^+ ions ✓ from the anode to the cathode. The Cl^- and OH^- ions are kept separate. This allows for a more

pure NaOH to form. ✓

- 7.1.5 Use the relative strengths of the oxidising agents present to explain why sodium metal is NOT produced at the cathode in the membrane cell (2)

Water has a higher oxidising ability than Na^+ , so water will be readily reduced at the cathode and not Na^+ to form Na metal.

- 7.1.6 A constant current of 4 000 A is passed through a chlor-alkali cell for 2,5 minutes. Calculate the volume of chlorine gas produced, at STP, at the anode during this period of time. (7)

$$Q = It = nF \quad \checkmark$$

$$4000 (2,5)(60) \quad \checkmark = n(96500) \quad \checkmark$$

$$n = 6,2176 \dots \text{mol}$$



$$1 : 2 \quad \checkmark$$

$$3,1088 \dots : 6,2176 \dots$$

$$V(\text{Cl}_2) = nV_m \quad \checkmark$$

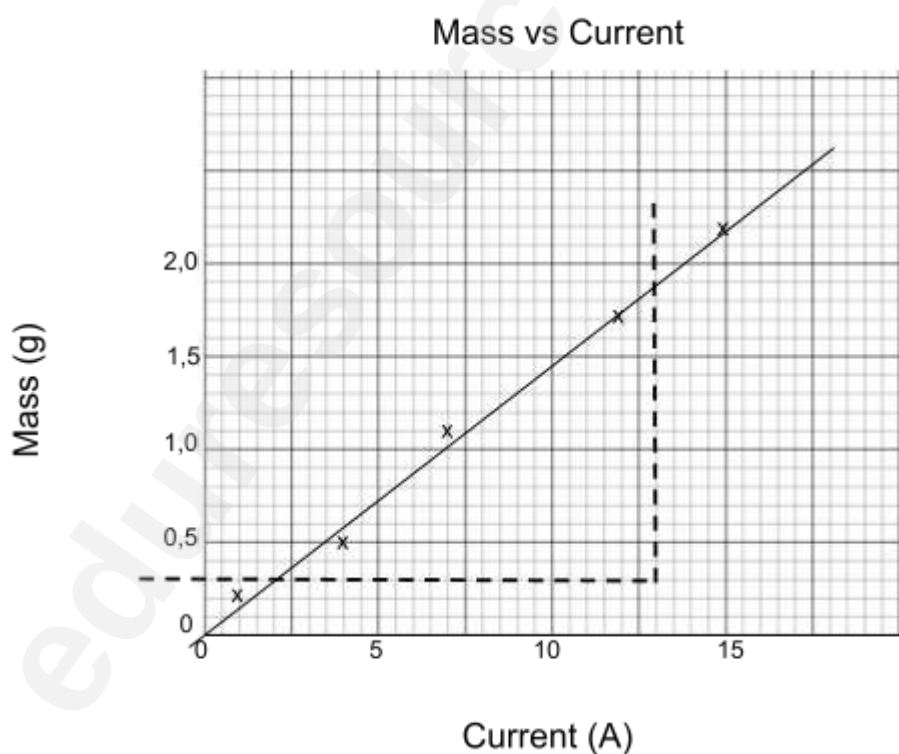
$$= 3,1088 \dots (22,4) \quad \checkmark$$

$$= 69,64 \text{ dm}^3 \quad \checkmark$$

- 7.2 In a trial done using a **mercury cell**, the effect of current on the mass of sodium produced at the cathode was measured. The current was passed through the brine for 10 minutes each time. The results are tabulated below.

Current (A)	Mass Na (g)
1	0,2
4	0,5
7	1,1
12	1,7
15	2,2

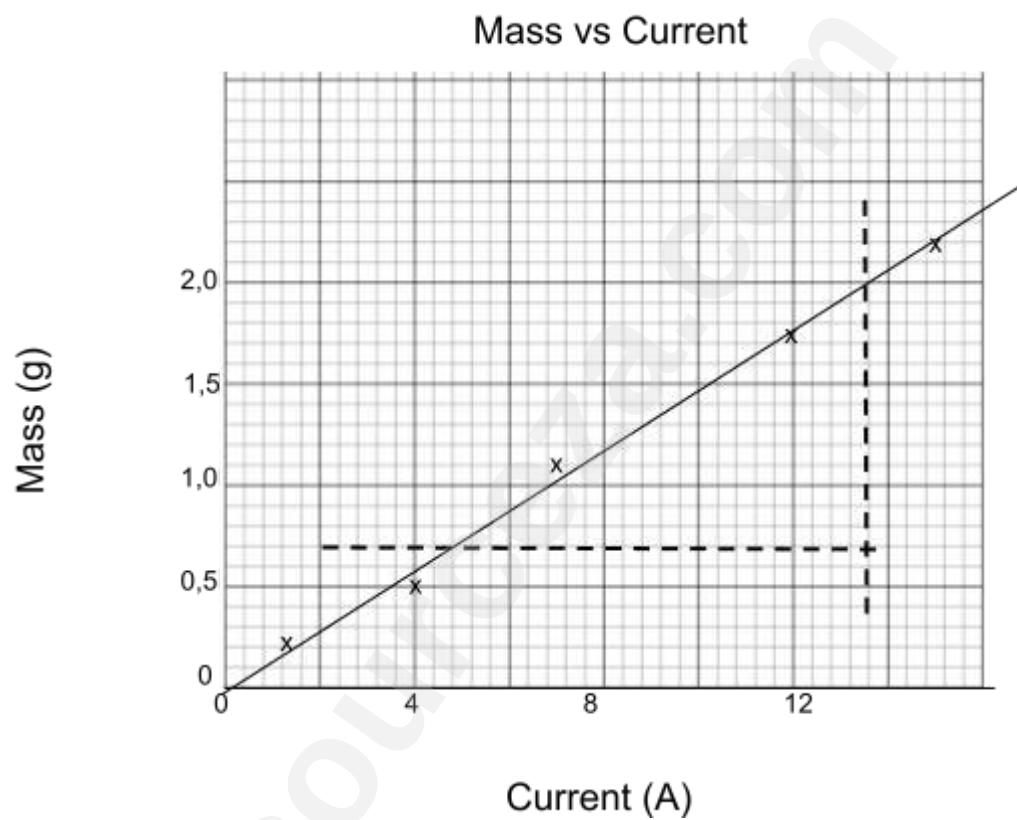
- 7.2.1 Plot a graph of the data with mass on the y-axis and current on the x-axis on the graph paper provided and draw a line of best fit through the points. (6)



- 7.2.2 Calculate the gradient of your graph. Show the points you used on your graph. (4)

$$\text{Gradient} = \frac{\Delta m}{\Delta I}$$

$$= \frac{1,8 - 0,3}{13 - 2} \checkmark \checkmark = 0,136 \text{ g.A}^{-1} \checkmark$$



- 7.2.3 Even though the mercury cell produced a relatively pure NaOH product, the use of the mercury cell has been discontinued. Discuss the advantages of using the Membrane Cell. (2)

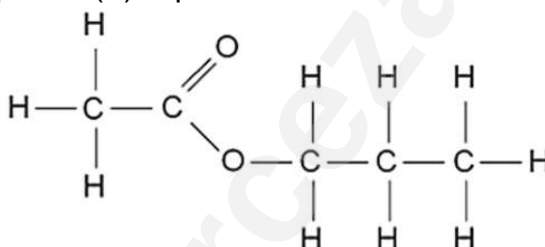
No mercury poisoning. ✓

Lower cost of electricity. ✓

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Question 8

The following compound (Z) is produced in a condensation reaction.



- 8.1 NAME the homologous series to which compound Z belongs. (1)

Ester ✓

- 8.2 Give the IUPAC name of compound Z. (2)

Propyl ethanoate ✓✓

- 8.3 Define *structural isomers*. (2)

Compounds having the same molecular formula but a different structural formula. ✓✓

- 8.4 Give the IUPAC name of the **functional isomer** of compound Z. (2)

Pentanoic acid ✓

- 8.5 Define a functional *group*. (2)

An atom or group of atoms that form the centre of chemical activity in the

molecule. ✓✓

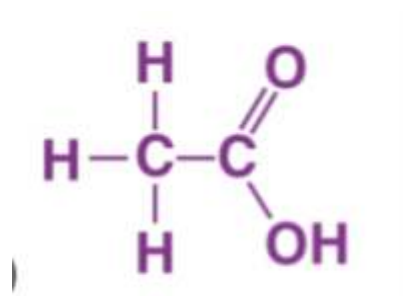
- 8.6 Give the NAME of the **functional group** of the isomer named in 8.4. (1)

Carboxyl group ✓

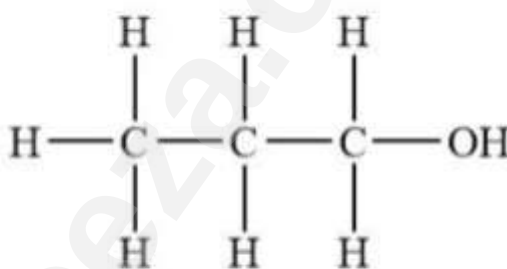
- 8.7 Give two reaction conditions must be met for the formation of compound Z. (2)

Heat and sulfuric acid ✓✓

- 8.8 Give the structural formulae of the two reactants needed for the formation of compound Z. (4)



Ethanoic acid

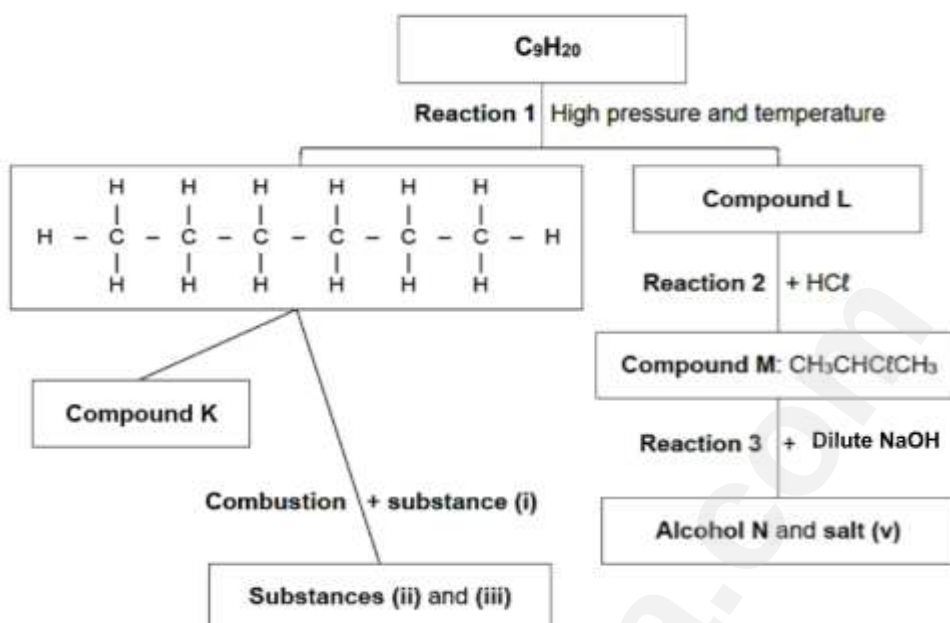


Propanol

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Question 9

A long chain hydrocarbon containing 9 carbons undergoes Reaction 1 to produce hexane and compound L. Further reactions of each substance are shown in the diagram below.



9.1 For Reaction 1, identify both the

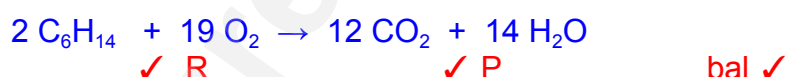
9.1.1 GENERAL type of reaction and (1)

Elimination ✓

9.1.2 SPECIFIC type of reaction that has occurred. (1)

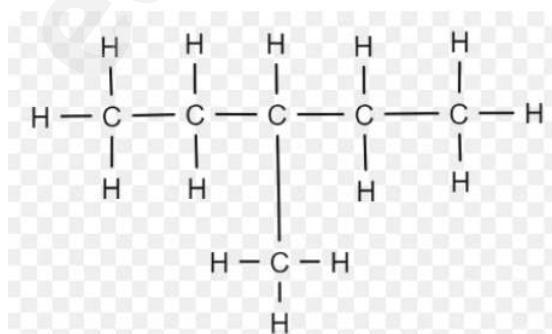
Cracking ✓

9.2 Write the balanced combustion reaction for hexane hence identifying substances (i)–(iii). Use molecular formulae. (Phase labels not required) (3)



9.3 Compound K is a structural isomer of hexane.

9.3.1 Draw the structural formula for a chain isomer of hexane. (2)



or any relevant isomer ✓ ✓

9.3.2 Give the IUPAC name for the isomer of hexane you have drawn. (2)

3-methylpentane ✓✓ (matching the isomer drawn in 9.3.1)

9.4 Give the IUPAC name for Compound L. (1)

Propene ✓

9.5 Refer to Reaction 2.

9.5.1 Identify the SPECIFIC type of reaction that took place. (1)

Hydrohalogenation ✓

9.5.2 What specific condition must be met for this reaction to take place? (2)

No water may be present ✓✓

9.6 Refer to Reaction 3 which produces an alcohol and a salt.

9.6.1 Identify the GENERAL type of reaction that took place. (1)

Substitution ✓

9.6.2 Give the molecular formula for alcohol N. (1)

C_3H_8O ✓

9.6.3 Name the salt. (1)

Sodium chloride ✓

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