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Definitions

1-Isotopes :

M1 atom(s) of the same element;
M2 with different number of neutrons;

2-Ionic Bond :

M1 (electrostatic) attraction;
M2 between oppositely charged ions;

3-Atomic number :

number of protons in one atom of an element;

4-Nucleon number :

M1 number of protons and neutrons in one atom of an element;
M2 in one atom of an element;

5- Catalyst :

M1 substance that speeds up a reaction/increases rate;
M2 unchanged (chemically) at the end/not used up/lowers activation energy/provides alternative pathway;

6- Electrolysis :

M1 breakdown of an ionic compound when molten or in aqueous solution;
M2 (using) electricity/electric current/electrical energy;

7- Unsaturated :

contains a double bond/triple bond/multiple bond;
OR
not all bonds are single bonds;

8-Proton number :

proton number: the number of protons

9- Amphoteric Oxide :

an amphoteric oxide will react with acids **AND** with bases

10- Neutral Oxide :

a neutral oxide will **not** react with acids **or** with bases

11-Hydrocarbon :

(compound that) contains carbon and hydrogen
and no other elements / only

12-Element :

(substance that) cannot be split up / broken down into (two or more) simpler substances by chemical means
OR
(substance) made of **atoms** with the same atomic number / number of protons / proton number

13- Compound :

(two or more) **elements** joined or combined or bonded (together)

14-Ion :

(particle) containing different numbers of protons and electrons
OR
atom or group of atoms that has gained or lost an electron / electrons

15- Saturated :

only single bonds **OR** no double / multiple bonds (between carbon atoms)

16- Structural Isomers :

(molecules with) the same molecular formula

different structural formula / different displayed formula

17- Polymer :

large / big molecule

made from (many) monomers (joined together)

18-Biodegradable :

(can be) broken down

by microbes / bacteria

19-Relative Atomic Mass :

average mass of (naturally occurring) atom(s) (of an element)
(compared to an atom of) ^{12}C

20-Excess :

more than enough to react

21-Strong Acid :

exists completely as ions (in solution) / completely dissociates (in solution) / completely ionises (in solution)

22- Condensation Polymerisation :

condensation:

M1 (two) molecules / monomers joining

M2 with the removal of a (small) molecule

polymerisation:

M3 (to form) a large molecule / a long chain

23- Weak Acid :

partially ionised / dissociated

24- Molecule :

(two or more) **atoms**

combined / joined / sharing electrons (by a covalent bond) / bonded

25- Equilibrium :

reversible reaction in which the **rate** of the forward reaction **equals** the **rate** of the backward reaction

concentration of all reactants and products becomes constant / does not change

26- Base :

proton / H^+ acceptor;

27- Thermal Decomposition :

M1 breaking down

M2 when heated

28- Ionic Lattice :

(ionic): made of, positive and negative ions / anions and cations / oppositely charged ions / unlike charged ions / different charged ions

(lattice): regular / sequence / pattern / alternating / repeated / framework / ordered / organised / network / uniform

29-Electrolyte :

substance that conducts electricity / (undergoes) electrolysis

decomposed / chemically changed

OR

molten or liquid or solution or aqueous **AND** containing ions/or ionic

30-Substitution Reaction :

(one) atom or group is replaced by another (atom or group)

31-Covalent Bond :

a shared **pair** of electrons (between two atoms)

M1 shared electrons

M2 pair of / two electrons

32- Mixture :

two (or more) substances not chemically combined

Good Luck 😊