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Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CHEMISTRY

0620/33

Paper 3 Core Theory

October/November 2016

MARK SCHEME

Maximum Mark: 80

Published

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[Turn over

Page 2	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2016	0620	33

Question	Answer	Marks
1(a)(i)	O / oxygen	1
1(a)(ii)	Li / lithium	1
1(a)(iii)	Cr / chromium	1
1(a)(iv)	Br / bromine	1
1(a)(v)	Ar / argon	1
1(b)	titanium lowest density strong / resistant to corrosion	1 1 1

Question	Answer	Marks
2(a)(i)	phosphate / PO_4^{3-}	1
2(a)(ii)	sulfate	1
2(a)(iii)	0.5 (g)	1
2(b)	<i>test:</i> aluminium / magnesium / Devarda's alloy sodium hydroxide / strong alkali (and warm) <i>result:</i> gas given off turns (red) litmus blue	1 1 1
2(c)(i)	filtration / filter	1
2(c)(ii)	carbohydrate AND protein	1
2(c)(iii)	random / zigzag / go anywhere / irregular	1

Page 3	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2016	0620	33

Question	Answer	Marks
2(d)(i)	any 2 from: - improve growth of plants - increase protein (in plants) - fertilisers add nitrogen / nitrates / phosphorous / phosphates / potassium - to put back nitrogen / nitrates / phosphorous / phosphates / potassium <u>into the soil</u>	2
2(d)(ii)	ammonia is produced / formed (ammonia) is a gas	1 1

Question	Answer	Marks
3(a)	<i>conditions required for ethanol manufacture by fermentation</i> (max = [3]) - uses yeast - uses glucose / sugar(s) - anaerobic / no oxygen present - room temperature / quoted temperature between 10 (°C)–40 (°C) (inclusive) - aqueous conditions / water needed - pH 7 / near pH 7 / neutral <i>conditions required for ethanol manufacture by hydration of ethene</i> (max = [3]) - uses high temperature / heat - uses a catalyst - uses high pressure - uses water / steam <i>equation</i> (max = [2]) - ethene + water / steam → ethanol - glucose → ethanol + carbon dioxide	5

Page 4	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2016	0620	33

Question	Answer	Marks
3(b)(i)	Liquid room temperature is between the melting point and boiling point (of methanol)/room temperature is above the melting point but below the boiling point (of methanol)	1 1
3(b)(ii)	values between 125(°C)–145(°C) inclusive	1
3(b)(iii)	increases with (increasing) number of carbon atoms	1
3(c)(i)	structure of ethanol showing all of the atoms and all of the bonds OH instead of O–H and rest of structure correct = [1]	2
3(c)(ii)	any suitable use, e.g. fuel/sterilisation/antiseptic solvent/making a named chemical, e.g. ethanoic acid/	1

Question	Answer	Marks
4(a)	any 3 from: • diffusion • particles move / motion of particles • (movement is) random / in any direction / in all directions • particles spread out / particles mix • particles move from high to low concentration	3
4(b)	<i>in pure water:</i> blue <i>in a strongly acidic solution:</i> yellow	1 1
4(c)(i)	A (volumetric) pipette B burette	1 1
4(c)(ii)	add (a few drops of) indicator to the flask slowly add acid (from the burette) into the alkali (until indicator) changes colour / until (alkali) neutralised / until neutral	1 1 1

Page 5	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2016	0620	33

Question	Answer	Marks
5(a)(i)	<u>endothermic</u> because heating is needed	1
5(a)(ii)	arrow(s) pointing in both directions / $\rightleftharpoons$	1
5(b)(i)	C in or just outside the tube at the top left L in or just outside the tube at the bottom right	1 1
5(b)(ii)	to produce a high temperature / for heat(ing)	1
5(c)	any 2 from: - plants / crops do not grow well if the soil is too acidic - increases the pH of the soil / makes the soil less acidic - neutralises the acid	2
5(d)(i)	<i>test:</i> (aqueous) barium chloride / (aqueous) barium nitrate <i>result:</i> white precipitate / white solid	1 1
5(d)(ii)	SiO_2 / Si_6O_{12}	1
5(e)(i)	pH 12	1
5(e)(ii)	H_2O	1
5(e)(iii)	any 3 from: (limewater absorbs) carbon dioxide (carbon dioxide) from the air - carbon dioxide dissolves in limewater - carbon dioxide (solution) is slightly acidic / carbon dioxide is an acidic oxide - idea that carbon dioxide reacts with / neutralises calcium hydroxide / neutralises limewater / neutralises the solution - pH (of limewater / solution) falls / pH goes down - calcium carbonate is formed	3

Page 6	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2016	0620	33

Question	Answer	Marks
6(a)(i)	(in order of their) atomic number / proton number	1
6(a)(ii)	less metallic across a period / metals on left of Periodic Table and non-metals on right of Periodic Table / electrical conductivity decreases / pattern in melting (or boiling) points (increases to a maximum then decreases)	1
6(a)(iii)	any 2 from: • density increases • melting / boiling point decreases • hardness decreases • reactivity increases	2
6(b)(i)	<i>from:</i> colourless / (light) green <i>to:</i> brown	1 1
6(b)(ii)	iodine is more reactive than astatine ORA	1
6(c)(i)	H ₂ (on left) 2(HCl) (on right)	1 1
6(c)(ii)	one pair of bonding electrons between H and Cl 6 non-bonding electrons around Cl and none around H	1 1
6(c)(iii)	lithium chloride water	1 1

Page 7	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2016	0620	33

Question	Answer	Marks
7(a)	A condensation / condensing / condense B freezing / solidification	1 1
7(b)	<i>arrangement:</i> regular <i>motion:</i> (only) vibrating / not moving (from place to place)	1 1
7(c)	<u>acidic</u> because phosphorous is a non-metal / phosphorous is on the right-hand side of the Periodic Table	1
7(d)	any 2 from: • does not conduct electricity / heat • has a low melting point / boiling point • insoluble in water / soluble in organic solvents	2
7(e)	sulfur dioxide is produced harmful effect of sulfur dioxide, e.g. acid rain / named effect of acid rain, e.g. corrodes metals / death of trees / kills organisms in lakes / irritation to lungs (or eyes / skin / nose / throat) /	1 1