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CAMBRIDGE INTERNATIONAL EXAMINATIONS

Cambridge International General Certificate of Secondary Education

MARK SCHEME for the October/November 2015 series

0620 CHEMISTRY

0620/21

Paper 2 (Core Theory), maximum raw mark 80

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Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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Page 2	Downloaded from eduresourceza.com Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2015	0620	21

Abbreviations used in the Mark Scheme

- ; separates marking points
- / separates alternatives within a marking point
- () the word or phrase in brackets is not required but sets the context
- **A** accept (a less than ideal answer which should be marked correct)
- **I** ignore (mark as if this material were not present)
- **R** reject
- ecf credit a correct statement that follows a previous wrong response
- ora or reverse argument
- owtte or words to that effect (accept other ways of expressing the same idea)

Page 3	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2015	0620	21

Question	Answer	Marks
1(a)(i)	F / barium chloride / BaCl_2 ;	1
1(a)(ii)	D / ammonium sulfate / $(\text{NH}_4)_2\text{SO}_4$;	1
1(a)(iii)	A / Al / Al_2Cl_6 / aluminium chloride;	1
1(a)(iv)	E / HCl / hydrogen chloride;	1
1(a)(v)	B / water / H_2O ;	1
1(a)(vi)	C / methane / CH_4 ;	1
1(b)(i)	arrow under the aluminium foil;	1
1(b)(ii)	fume cupboard;	1
1(b)(iii)	3 (Cl_2);	1

Question	Answer	Marks
2(a)	make sure temperature change is the same throughout / make sure that there are no hot spots / no local heating;	1
2(b)	any two from: • same amount of solid / same mass of solid; • same volume of water; • same amount of stirring;	2
2(c)(i)	Q	1
2(c)(ii)	R; T;	2
2(d)(i)	^{235}U ;	1
2(d)(ii)	138;	1
2(d)(iii)	cancer treatment / tracer / test thyroid function;	1
2(e)(i)	<u>kerosene</u> ;	1
2(e)(ii)	C_5H_{10} ;	1

Page 4	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2015	0620	21

Question	Answer	Marks
3(a)	2 (Ni); CO ₂ (on right);	1 1
3(b)(i)	positive electrode / anode in box on left; negative electrode / cathode in upper box on right; electrolyte / named suitable electrolyte in lower box on right; 3 correct = [2] 1 or 2 correct = [1]	1 1
3(b)(ii)	cathode / negative electrode;	2
3(c)(i)	<i>the positive electrode</i> : chlorine; <i>the negative electrode</i> : nickel;	1
3(c)(ii)	inert / unreactive; conducts electricity;	1 1
3(d)(i)	giant structure / lots of carbon atoms joined to each other / lattice of covalent bonds; strong (covalent) bonds throughout;	1 1
3(d)(ii)	weak forces between layers; layers can slide (over each other);	1 1

Question	Answer	Marks
4(a)(i)	gradient / slope is greater for strontium ora;	1

Page 5	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2015	0620	21

Question	Answer	Marks
4(a)(ii)	11 (cm ³);	1
4(a)(iii)	64–66 (s);	1
4(a)(iv)	the line was still going up / the line was still rising;	1
4(a)(v)	(rate) increases;	1
4(b)(i)	(volumetric) pipette;	1
4(b)(ii)	to show end point of titration / to show when the solution has been neutralised; litmus goes from blue to pink (at end point);	1 1
4(c)(i)	decreases slowly at first; then sudden decrease in pH; then slow decrease;	1 1 1
4(c)(ii)	26 (cm ³);	1
4(c)(iii)	strontium chloride;	1

Page 6	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2015	0620	21

Question	Answer	Marks
5(a)	melting/ice melts/ice goes from solid to liquid;	1
	any four from: • in solid particles regularly arranged; • in solid particles arranged in fixed position / cannot move; • particles in solid absorb energy; • particles (in solid) vibrate more / particles start to move when heated; • forces between particles (in solid) broken; • particles in liquid slide over each other / move; • particles in liquid not regularly arranged;	4

Page 7	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2015	0620	21

Question	Answer	Marks
5(b)(i)	coolant / for making ethanol / for making specified chemicals / solvent;	1
5(b)(ii)	washing / cooking / cleaning etc.;	1
5(c)(i)	lithium + water → lithium hydroxide + hydrogen;	1
5(c)(ii)	any two from: • floats on surface (of water); • bubbles of gas; • fizzes / fizzing sound; • decreases in size / disappears; • moves around;	2
5(c)(iii)	(potassium) more reactive / lithium less reactive;	1
5(d)(i)	correct structure of ethene;	1
5(d)(ii)	high temperature; catalyst/phosphoric acid;	1 1
5(e)	cobalt → iron → cerium → calcium; one pair reversed or all reversed = [1]	2

Question	Answer	Marks
6(a)	C ₅ H ₈ ;	1

Page 8	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2015	0620	21

Question	Answer	Marks
6(b)	bromine / bromine water / aqueous bromine; decolourised / goes from orange to colourless;	1 1
6(c)(i)	double bond;	1
6(c)(ii)	poly(ethene) / any other addition polymer;	1
6(d)	it is a molecule / covalent compound;	1
6(e)	any two from: carbon or soot / carbon monoxide / water;	2
6(f)	<u>alcohols</u> / first box ticked;	1

Question	Answer	Marks
7(a)	any five from: • 11 electrons; • electrons – (negatively) charged; • electrons outside nucleus in shells; • nucleus contains protons and neutrons; • protons – (positively) charged; • neutrons no charge; • 11 protons; • 12 neutrons; • electron arrangement 2,8,1 / 1 electron in outer shell;	5
7(b)(i)	2 (NaOH); 2 (H ₂ O);	1 1
7(b)(ii)	substance containing carbon and hydrogen <u>only</u> / substance containing carbon and hydrogen and <u>no other</u> element;	1
7(b)(iii)	70;	1